

5.2a Naming Acids

Science 10 Notes

Naming Acids

Acids have common names:

HCl (stomach acid) : _____

H₂SO₄ (in car batteries) : _____

CH₃COOH (vinegar) : _____

Where do these names come from?

Acids are named based on their chemical formulas. There are two decisions to be made:

1. Does the acid begin with "hydro" or not?
2. Does the acid end in -ic or -ous

Hydro -ic acids

Look at how the chemical formula would be named based on the naming rules we developed last chapter.

- If the anion (negative ion) would be named and end with _____, then the acid name starts with _____ and it will end with _____

No Hydro, -ic acids

- If the anion is a polyatomic ion whose name ends in _____, then the name does not begin with _____ and the name will end with _____.

No Hydro, -ous acids

- If the anion is a polyatomic ion whose name ends in _____, then the name does not begin with _____ and the name will end with _____.

5.2b Strength of Acids and Bases

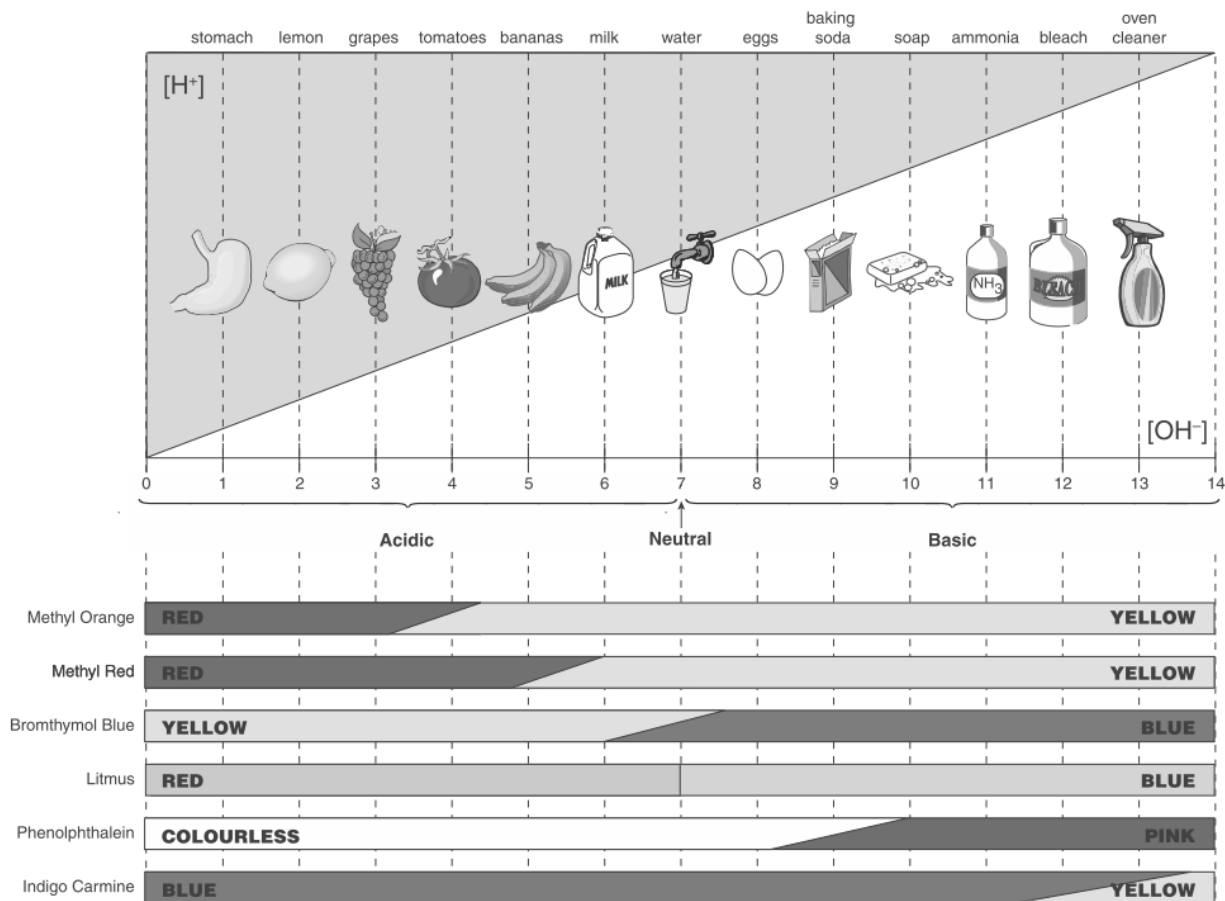
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pH Scale

Strengths of Acids and Bases are measured on the same scale, called the _____ scale.

Each unit on the scale shows an increase in strength by a factor of _____.

pH measures the relative amounts of _____ and _____ ions.



Indicators

Indicators don't actually tell us that something is an acid or base, but the _____ where they change colour can tell us whether an unknown solution is an acid or a base.