

Science 10

Notes: Formulas of Covalent Compounds

how to name a compound
how to write formula if you know name.

Review:

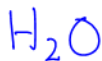
Covalent compounds are formed between non-metal and non-metal

The compound is formed because: both non-metals want to take electrons, so they end up sharing.

Naming Compounds

Rules:

- left element is said first
- The right element is given the suffix "-ide"
- The number of atoms of each type determines what prefix to attach to the molecule **unless** hydrogen is involved.
 - The first element only gets a prefix if there are more than one
 - The second element always gets a prefix
 - Don't follow the "lowest terms" rule for compounds



monoxide

hydrogen monoxide
or

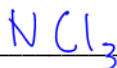
dihydrogen monoxide

Note: This is only done for covalent compounds and not for ionic compounds

Prefixes used in naming	
Prefix	Number
mono-	1
di-	2
tri-	3
tetra-	4
penta-	5
hexa-	6
hepta-	7
octa-	8
nona-	9
deca-	10

Write the formulas for each compound

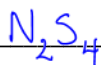
Nitrogen trichloride



Sulfur hexafluoride



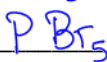
Dinitrogen tetrasulfide



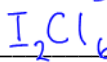
Sulfur trioxide



Phosphorous pentabromide



Diiodine hexachloride



Write the name for each compound

N_2O

dinitrogen monoxide

CO_2

carbon dioxide

PI_3

phosphorous tri-iodide

N_2O_4

dinitrogen tetraoxide.

PCl_5

phosphorous pentachloride.

P_4S_{10}

tetraphosphorous decasulfide.

NO

nitrogen monoxide.

workbook
p71-72.

