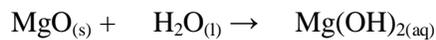
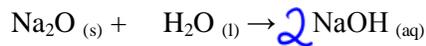


Oxides

These are special compounds that react to form acids and bases

Metal Oxides

- React with water to form bases



already balanced

(s) = solid

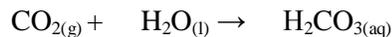
(l) = liquid

(aq) = dissolved in water

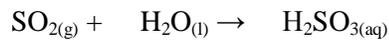
(g) = gas.

Non-Metal Oxides

- React with water to form acids



→ soda pop fizz.



non Metal oxides are often produced by burning fossil fuels
These are of great concern because they react with atmospheric water to form acid rain

P85-86

p93-94.

Organic Compounds

Organic chemistry is the study of compounds that contain carbon

Almost all carbon-containing compounds are organic (including both natural and synthetic)

Inorganic compounds generally **do not** contain carbon

exceptions:

CO₂
CO

and anything with CO₃²⁻ is inorganic

Carbon is very versatile due to having 4 valence electrons, and can bond to H, O or

other non metals

Carbon can also bond to itself as well, and can form long chains of carbon atoms called polymers. These are how plastic and petroleum are formed.

Formulas for organic compounds will always have C before H. Eg (CH₄) methane

Some organic compounds:

Hydrocarbons – contain only C and H.

- Highly flammable, and often used as fuels eg CH₄ (methane) - heats homes
C₃H₈ (propane) - BBQ

Alcohols – contain only C, H and O.

- Flammable and poisonous.

Carbohydrates – contain only C, H, O.

- Important compound for all living things

— sugars
— starches
— wood (gives structure to plants)

p 85-86 p 93-94 acids, oxides ← finish for HW
read p 95-97 do p 98-99 organics ←

finished for
next Monday.